









<sup>6</sup>  
Constant-Pressure Lines on the T-s Diagram  
Gibbs  
Equation  
A perfect gas 
$$Tds = du + pdv$$
  
A perfect gas  $pv = RT$   
 $pdv + vdp = RdT$   
 $du = C_v dT$   $dh = C_p dT$   $C_p = C_v + R$   
 $Tds = C_v dT + RdT - vdp = C_p dT - vdp$   
Thus:  $\frac{\delta T}{\delta s_p} = \frac{T}{C_p}$ 











































































































